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Ionic Compounds With Polyatomic Ions, Part 4

Get to zero!

+ 3 - 1 - 1 - 1 = 0
+ 3 + 3 - 2 - 2 - 2 = 0

A.	Al^{+3}	NO_3^{-1} NO_3^{-1} NO_3^{-1}	$Al(NO_3)_3$	Aluminum Nitrate
	Ga^{+3}	NO_2^{-1} NO_2^{-1} NO_2^{-1}	$Ga(NO_2)_3$	Gallium Nitrite
	In^{+3}	ClO^{-1} ClO^{-1} ClO^{-1}	$In(ClO)_3$	Indium Hypochlorite
	In^{+3}	ClO_2^{-1} ClO_2^{-1} ClO_2^{-1}	$In(ClO_2)_3$	Indium Chlorite
	Y^{+3}	ClO_3^{-1} ClO_3^{-1} ClO_3^{-1}	$Y(ClO_3)_3$	Yttrium Chlorate

B.	Al^{+3}	Al^{+3} SO_4^{-2} SO_4^{-2} SO_4^{-2}	$Al_2(SO_4)_3$	Aluminum Sulfate
	Ga^{+3}	Ga^{+3} SO_3^{-2} SO_3^{-2} SO_3^{-2}	$Ga_2(SO_3)_3$	Gallium Sulfite
	In^{+3}	In^{+3} CO_3^{-2} CO_3^{-2} CO_3^{-2}	$In_2(CO_3)_3$	Indium Carbonate
	In^{+3}	In^{+3} O_2^{-2} O_2^{-2} O_2^{-2}	$In_2(O_2)_3$	Indium Peroxide
	Y^{+3}	Y^{+3} CrO_4^{-2} CrO_4^{-2} CrO_4^{-2}	$Y_2(CrO_4)_3$	Yttrium Chromate

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